**Name: Date:**

**Solubility of Salt Lab**

**Post-Lab Questions:** Answer these questions using as much information from the lab as possible. Then, use your answers to help guide your “**Discussion**” in your formal lab report.

1. Solve for molarity of your beakers of salt water at the end of Part A:
2. Describe the general trend (if any) that you observed of the relationship between temperature and solubility of salt. How do you explain this observation?
3. Solve for the rate of dissolving of 1.00 g of salt at each of the following temperatures. Your rate should be in (**mg/sec**). Equation: $\frac{mass of salt added (mg)}{dissolve time (sec)}$
	1. Ice Water:
	2. Room temperature:
	3. 40 °C:
	4. 60 °C:
	5. 80 °C:
4. Describe the relationship between temperature and rate of dissolving.
5. Temperature seems to be an important factor in this lab. What about temperature have we discussed in class that might explain its affects on solubility and the rate of dissolving?