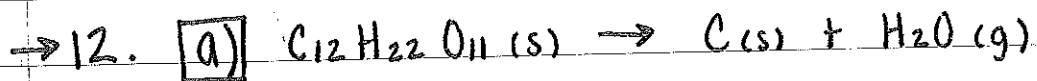
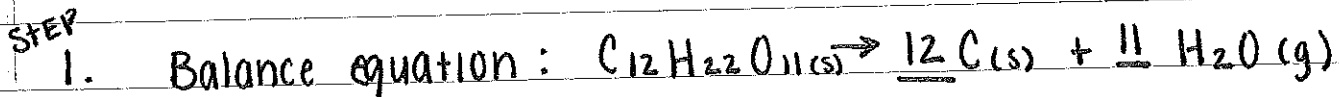


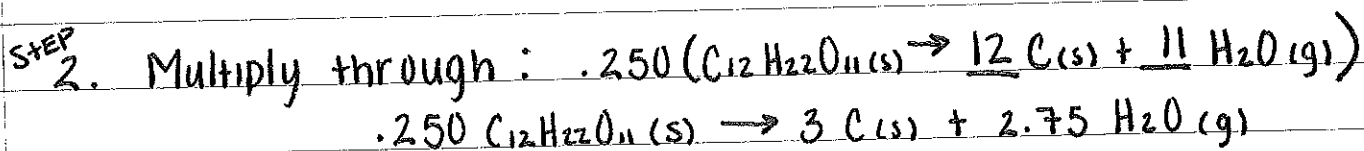
# Homework Answers 2/07/12



STEP

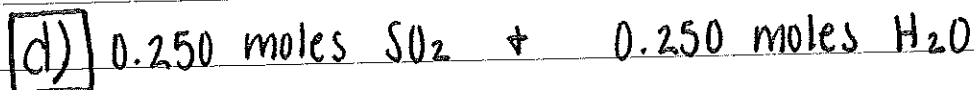


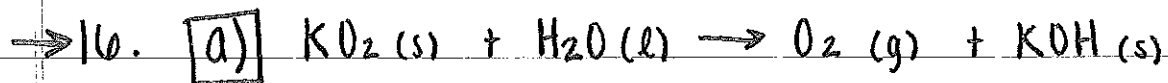
STEP



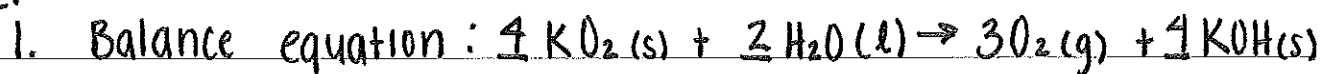
STEP

3. Clarify answer: .250 moles of  $C_{12}H_{22}O_{11}$  will form  
 $\therefore 3.00$  moles of C +  $2.75$  moles  $H_2O$

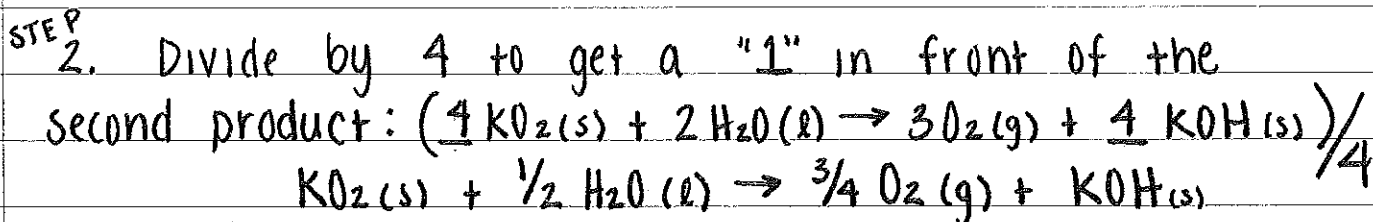




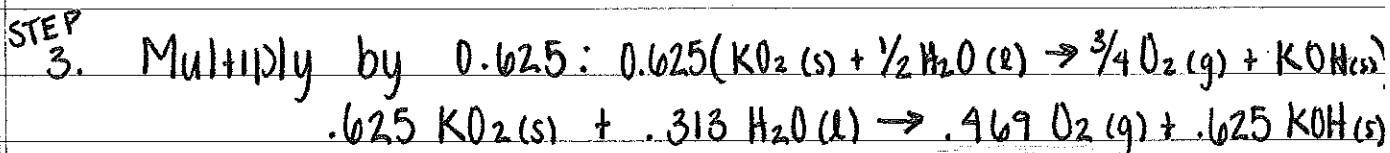
STEP



STEP



STEP



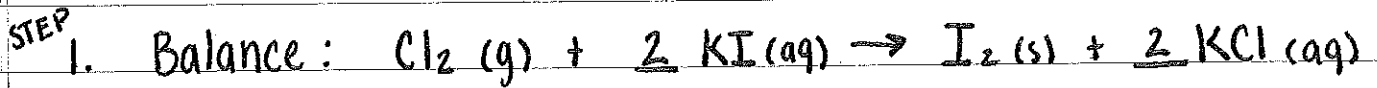
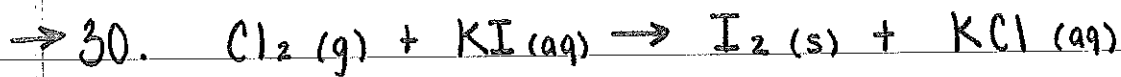
STEP

4. Clarify answer: .625 moles of KOH will form  
 $\therefore .469$  moles  $\text{O}_2$

**b)** 0.938 mol Se

**c)** 0.625 mol  $\text{CH}_3\text{CHO}$

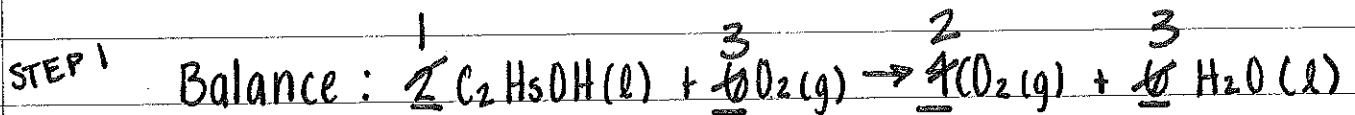
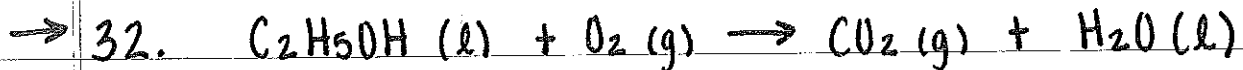
**d)** 1.25 mol Fe



STEP 2. Solve:

$$2.55 \text{g Cl}_2 \times \frac{1 \text{ mol Cl}_2}{70.90 \text{g Cl}_2} \times \frac{1 \text{ mol I}_2}{1 \text{ mol Cl}_2} \times \frac{253.8 \text{g I}_2}{1 \text{ mol I}_2} = ?$$

$$\therefore 9.13 \text{ g I}_2$$



STEP 2. Solve:

$$25.0 \text{g C}_2\text{H}_5\text{OH} \times \frac{1 \text{ mol C}_2\text{H}_5\text{OH}}{46.07 \text{g C}_2\text{H}_5\text{OH}} \times \frac{4 \text{ mol CO}_2}{2 \text{ mol C}_2\text{H}_5\text{OH}} \times \frac{44.01 \text{g CO}_2}{1 \text{ mol CO}_2} =$$

$$\therefore 47.8 \text{ g CO}_2$$