

Name \_\_\_\_\_

Remember: atoms & molecules  
can both be used in  
conv fact:  $\frac{1 \text{ mole}}{6.02 \times 10^{23} \text{ atoms or molec.}}$

Moles, Molecules, and Grams Worksheet

- 1) How many molecules are there in 24 grams of  $\text{FeF}_3$ ?
  
  
  
  
  
  
  
  
  
  
- 2) How many molecules are there in 450 grams of  $\text{Na}_2\text{SO}_4$ ?
  
  
  
  
  
  
  
  
  
  
- 3) How many grams are there in  $2.3 \times 10^{24}$  atoms of silver?
  
  
  
  
  
  
  
  
  
  
- 4) How many grams are there in  $7.4 \times 10^{23}$  molecules of  $\text{AgNO}_3$ ?
  
  
  
  
  
  
  
  
  
  
- 5) How many grams are there in  $7.5 \times 10^{23}$  molecules of  $\text{H}_2\text{SO}_4$ ?
  
  
  
  
  
  
  
  
  
  
- 6) How many molecules are there in 122 grams of  $\text{Cu}(\text{NO}_3)_2$ ?
  
  
  
  
  
  
  
  
  
  
- 7) How many grams are there in  $9.4 \times 10^{25}$  molecules of  $\text{H}_2$ ?
  
  
  
  
  
  
  
  
  
  
- 8) How many molecules are there in 230 grams of  $\text{CoCl}_2$ ?