**Ch. 9.5 Lecture Guide**

**H. Chemistry**

9.5

* **Theoretical Yield**
* **Actual Yield**
* **Percent Yield**
	+ **Consider the following problem:** Suppose 6.85 x 104 g of carbon monoxide is reacted with 8.60 x 103 g of hydrogen gas to produce methanol (CH3OH).
		- **(a) Calculate the theoretical yield of methanol.** **(b) If 3.57 x 104 g of methanol is actually produced, what is the percent yield of methanol?**
			* **(a)**
			* **(b)**