**Name:\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ Date:\_\_\_\_\_\_\_**

**H. Chemistry**

**Ch. 8.1-8.2 Lecture Guide**

* Candy Shop:
  + Food for thought: Write down a possible solution to this father’s dilemma. How might he make sure that his daughters get the same number of pieces of candy?
* Atomic Masses: Counting Atoms by Weighing
  + Atomic Mass Unit:
    - 1 amu = 1.66 x 10-24g
  + To find the average atomic mass of an atom, look at the number below an element’s symbol on the periodic table.

|  |  |
| --- | --- |
| **Element** | **Average Atomic Mass (amu)** |
| Hydrogen |  |
| Carbon |  |
| Nitrogen |  |
| Oxygen |  |
| Sodium |  |

* Practice Problems
  + Calculate the mass, in amu, of a sample of 300 oxygen atoms.
  + Calculate the number of sodium atoms present in a sample that has a mass of 1172.49 amu.